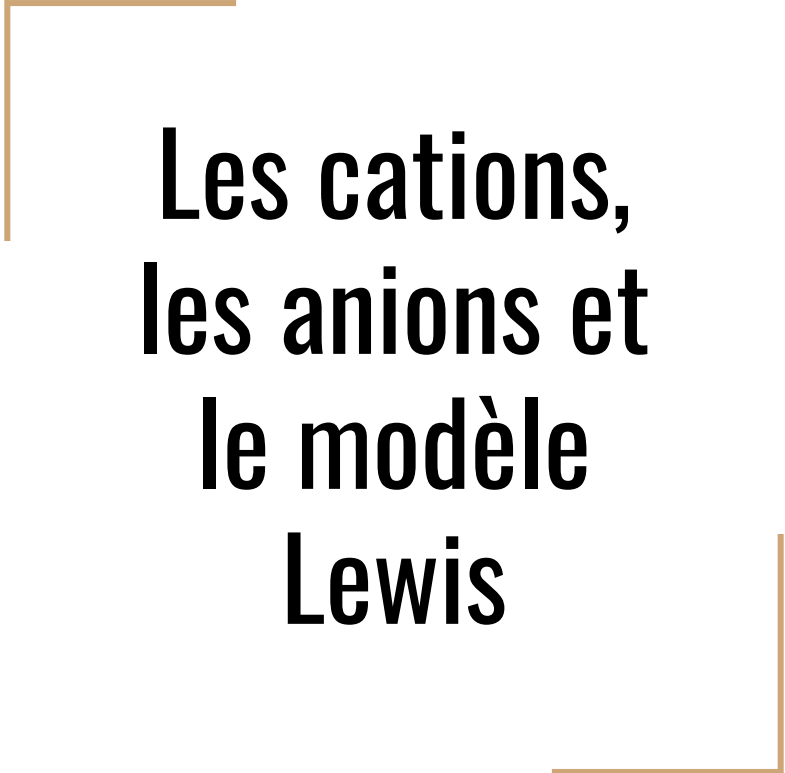


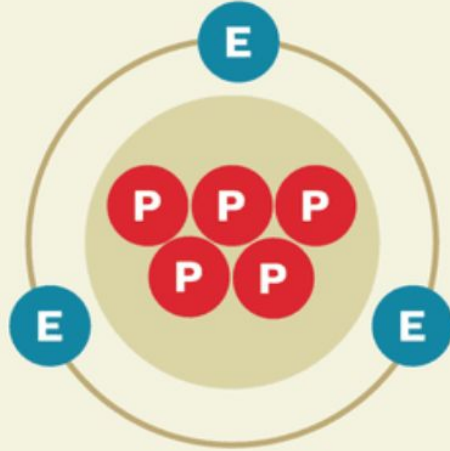
1.5 Ions et Lewis

Oct 22, 2018 at 13:39



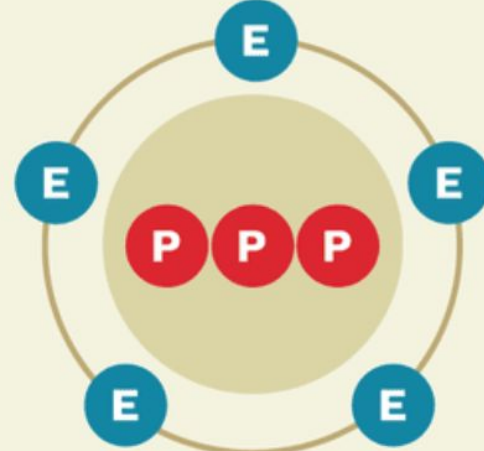
**Les cations,
les anions et
le modèle
Lewis**

CATION



- ❑ Charge positif
- ❑ Plus de protons que d'électrons

ANION



- ❑ Charge négatif
- ❑ Plus d'électrons que de protons

L'atome

Un atome peut être en deux formes:

- ❑ Neutre Al
- ❑ Ion Al^{3+}

Un atome devient un ion lorsqu'il réagit avec un autre atome. Il réagit afin d'avoir une couche de valence remplie.

Si un atome devient un *anion* ou un *cation* dépend de l'élément.

Surtout sur le # d'électrons de valence.

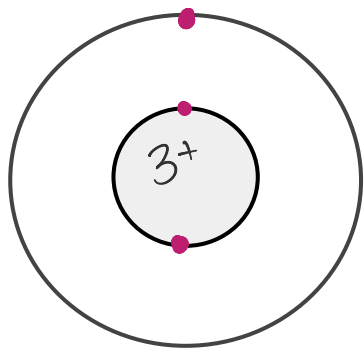
Lithium

$P = 3$

$E = 3$

$N = 4$

Atome neutre



Ion

Li va perdre $1e^-$

3 protons

2 électrons

\bullet
Li



Charge: $+1$

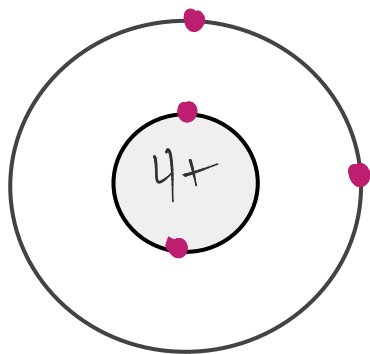
Béryllium

$$P = 4$$

$$E = 4$$

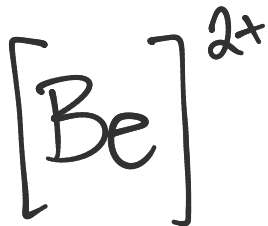
$$N = 5$$

Atome neutre



Ion

Be va perdre 2 e⁻



Charge:



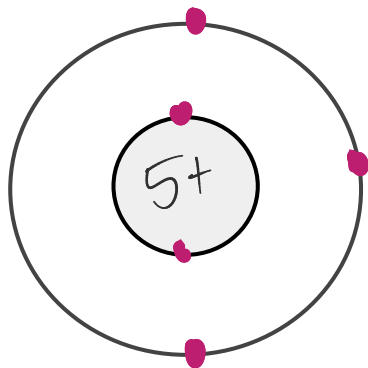
Bore

$$P = 5$$

$$E = 5$$

$$N = 6$$

Atome neutre



Ion

B va perdre $3 e^-$

5 Protons

2 électrons

Charge: 3^+

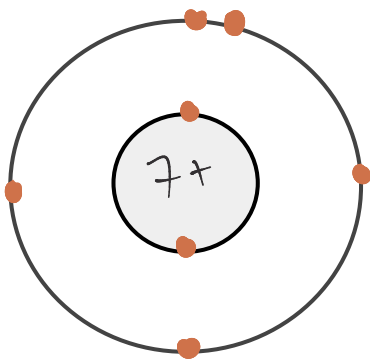
Azote

$$P = 7$$

$$E = 7$$

$$N = 7$$

Atome neutre

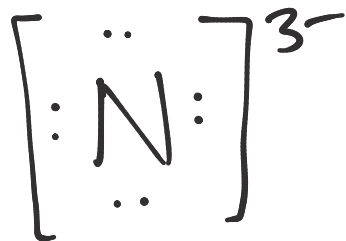
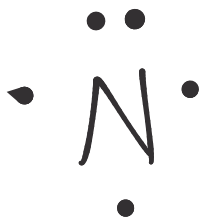


Ion

N va gagner $3e^-$

7 protons

10 électrons

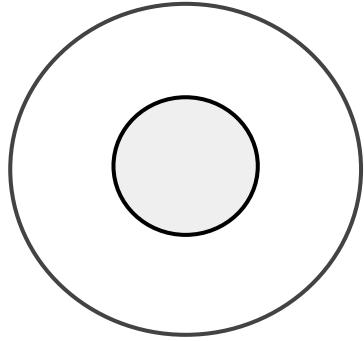


Charge:



Oxygène

Atome neutre

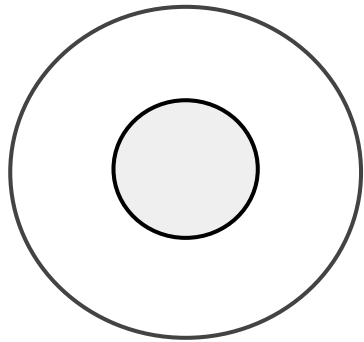


Ion

Charge:

Fluor

Atome neutre



Ion

Charge:

1+

2+



1

1

1

1

1

1

1

2

2

2

2

2

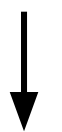
2

3+

3-

2-

1-



2

3

4

5

6

7

8

3

4

5

6

7

8

3

4

5

6

7

8

3

4

5

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7

8

3

4

5

6

Détermine le PEN atomique et dessine la structure Bohr

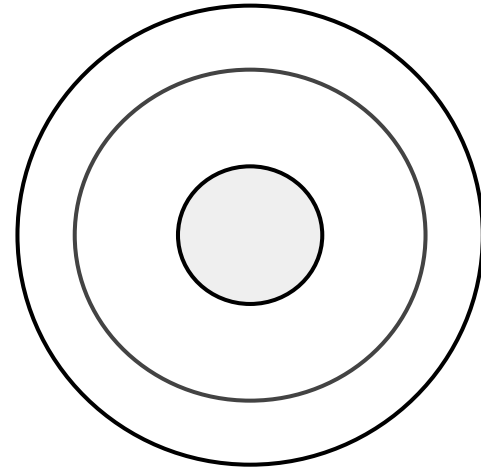
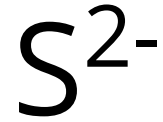
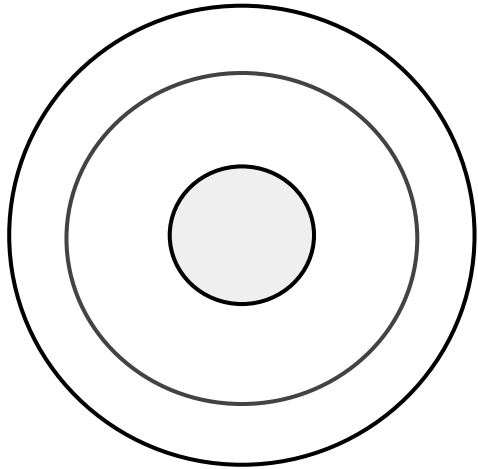
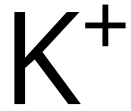


Protons 20^+

Électrons $18e^-$

Neutrons 20

Détermine le PEN atomique et dessine la structure Bohr



	Élément	# d'électrons de valence	Gagne ou perd ____ e ⁻	Charge de l'ion
1	Cl			
2	Na			
3	Mg			
4	O			
5	N			
6	Al			
7	Xe			

